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A. 1.20 10<sup>×</sup> 23 B. 3.01 10<sup>×</sup> 23 C. 6.02 10<sup>×</sup> 23 D. 1.20 10<sup>×</sup> 24 2. 1 mol of a hydrocarbon with general formula C<sub>n</sub>H<sub>2n+2</sub> reacts completely with oxygen to produce 4 mol of H<sub>2</sub>O. What is the amount of oxygen molecules, in mol, that reacts? A. 4 B. 5 C. 6 D. 7 3. Under which combination of conditions is 1 mol of an ideal gas present? Volume Pressure ...

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The molar mass of a compound is approximately 56 g mol<sup>-1</sup>. Which formula is possible for this  
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